

Name: _____ Date: _____ Period: _____

AP Chemistry Chapter 13 Essentials Pt II

EVAPORATION

1. Why can a liquid change phase to a gas at temperatures below its boiling point?
2. What is a dynamic equilibrium and what principle does it illustrate?
3. Describe what is meant by the vapor pressure of a liquid.
4. Explain how the volatility of liquids can be explained by the intermolecular forces between molecules. Use the molecules in figure 13-13 to explain your answer.
5. How is the process of boiling different from evaporation/vaporization?

6. How can distillation be used to separate out a mixture of liquids?

7. How does altitude and pressure affect the boiling point temperature of a liquid like water?

8. How are the molar heat capacity and enthalpy of vaporization used to determine heat transfer in liquids?

9. What is the specific heat and enthalpy of vaporization for water?